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Solution . To get the requested answer, let us formulate a "stoichiometric" equation (molar quantities) for the reaction: C H + 70 2CO + 3H O. 26 → 2 2 Each C. 2H 6 (ethane) molecule requires 7 oxygen atoms for complete combustion. In molar quantities: 1 mole of C. 2H 6 = 2 x 12.01 + 6 x 1.008 = 30.07 g requires . 7 x 15.9984 g = 1.12 x 10

Session #2: Homework Solutions

Lesson #8 – Introduction to Logarithms Lesson #9 – Graphs of Logarithms Lesson #10 – Logarithm Laws Lesson #11 – Solving Exponential Equations Using Logarithms Lesson #12 – The Number e and the Natural Logarithm Lesson #13 – Compound Interest Lesson #14 – Newton's Law of Cooling

COMMON CORE ALGEBRA II

1 Introduction 1 Motivation to Study Combustion 1 A Definition of Combustion 6 Combustion Modes and Flame Types 6 Approach to Our Study 8 References 8 2 Combustion and Thermochemistry 9 Overview 9 Review of Property Relations 9 Extensive and Intensive Properties 9 Equation of State , 10 Calorific Equations of State 11 Ideal-Gas Mixtures 13

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