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SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio SECTION 1 Defining Stoichiometry 368

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Answer the following questions in the space provided. 1. bThe coefficients in a chemical equation represent the. (a)masses in grams of all reactants and products. (b)relative number of moles of reactants and products. (c)number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW

Stoichiometry Worksheet #1 Answers 1. Given the following equation: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$, show what the following molar ratios should be. a. $\text{C}_4\text{H}_{10} / \text{O}_2$ b. O_2 / CO_2 c. $\text{O}_2 / \text{H}_2\text{O}$ d. $\text{C}_4\text{H}_{10} / \text{CO}_2$ e.

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Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a masses in grams of all reactants and products. (h) relative number of moles of reactants and products.

Date. FCHAPJ REV[EW.

1 mol Al 2 O 3 or 1 mol Al 2 O 3 ___
101.96 g Al 2 O 3 26.98 g Al _ 1 mol Al
or _ 1 mol Al 26.98 g Al 32.00 g O 2 _ 1
mol O 2 or 1 mol O 2 _ 32.00 g O 2 To
find the number of grams of aluminum
equivalent to 26.0 mol of aluminum, the
calculation would be as follows. 26.0 mol
Al \times 26.98 g Al / 1 mol Al = 701 g Al
Stoichiometry 291 SECTION 1 ...

CorrectionKey=NL-A DO NOT

EDIT--Changes must be made ...

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CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H

CHAPTER 9 REVIEW Stoichiometry

Chapter 12 Stoichiometry 127 SECTION

12.1 THE ARITHMETIC OF EQUATIONS

(pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Chapter 14 and 15 Acid Base Chemistry

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